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A compound analyzed in a chemistry lab consists of 5.34 g of carbon, 0.42 g of Hydrogen, and 47.08 of Chlorine.

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Chemical Quantities.

STUDY. PLAY. Mole.

(mol) of a substance is  
6.02 times  $10^{23}$

representative

particles of that

substance and is the SI

unit for measuring the

amount of a substance.

Avogadro's Number.

the number of

representative

particles in a mole,

6.02 times  $10^{23}$ .

Representative

Particle.

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STUDY. PLAY. mole.

$6.02 \times 10^{23}$

representative

particles of a

substance, SI unit for

measuring the amount

of a substance.

Avogadro's number.

$6.02 \times 10^{23}$

representative

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particles of a substance, named in honor of Italian scientist Amedeo Avogadro di Quaregna.

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Chapter 10. Avogadro's  
number. Empirical  
Formula. Molar Mass.  
Molar Volume. number  
of particles in one mole  
of a pure substance  
(element o.... Formula  
that shows the lowest

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whole-number ratio of the atoms... The mass of one mole of an element. Found on the periodic tabl....

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...

10.2 Mole to Mole &  
Mole to Volume  
Relationships \*For all  
the problems on this  
page, first find the  
molar mass of the  
compound. 1. What is  
the mass of 9.45 mol of  
aluminum oxide? mol A  
ILO + 3 (I A GO q  
63,5ù?- 2. What is the

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### Chapter 10

mass of  $4.52 \times 10^{-3}$  mol  
of ethylbenzene,  
 $C_6H_5CH_2CH_3$ ?

\*Ethylbenzene is a hydrocarbon that is produced by burning coal.

### **BHS - Moodle**

A chemistry student working in the lab might be asked to calculate how much 1-bromo-2-methylpropane,  $C_4H_9Br$ , could be made from 6.034 g of ... 10 1 kg 103 g 368

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Chapter 10 Chemical  
Calculations and  
Chemical Equations.  
10.1 Equation

Stoichiometry 369 The  
ratio of moles of P 4O

## **Chapter 10 Chemical Calculations and equations**

6.7 Chapter Summary.  
To ensure that you  
understand the  
material in this  
chapter, you should  
review the meanings of  
the following bold

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terms in the following summary and ask yourself how they relate to the topics in the chapter. Chemical reactions relate quantities of reactants and products.

### **Chapter 6 - Quantities in Chemical Reactions - Chemistry**

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Quantities. 10.1 The  
Mole: A Measurement  
of Matter - Sample

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Problem 10.1. 10.1 The Mole: A Measurement of Matter - Chemistry & You. 10.1 The Mole: A Measurement of Matter - Sample Problem 10.2. 10.1 The Mole: A Measurement of Matter - Sample Problem 10.3.

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...

Chapter 10 "Chemical  
Quantities" Vocab. the  
SI unit representing  
 $6.02 \times 10^{23}$   
representative  
particles of a  
substance. the  
temperature and  
pressure at which one  
mole of gas occupies a  
volume of 22.4 L. equal  
volumes of gases at

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the same temperature and pressure contain equal numbers of particles.

### **Quia - Chapter 10 "Chemical Quantities" Vocab**

Use the chemical formula to find the number of atoms in one molecule and multiply this number by Avogadro's number, the number of particles in one mole.

$$\bullet 1023 \text{ O } 2 \text{ } 6.02 \bullet 10$$

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23 ion  $\text{Na}^+$  6.02 •

1023 formula unit  $\text{NaCl}$

6.02 • 1023 6.02 •

1023 representative  
particles of a  
substance molecule  
formula unit atom

**Chemical Quantities**

**- Mr. Mutic's**

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Answers substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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SECTION 10.1 THE ...

Quantities

## **Chapter 10 Chemical Quantities Guided Practice Answers**

Here is an example.

Pure HCl (hydrogen chloride) is a gas that is very soluble in water. A plot of the partial pressure of gaseous HCl in equilibrium with aqueous HCl, as a function of the solution molality (Fig. 10.1),

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shows that the limiting slope at infinite dilution is not finite, but zero.

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