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Chemistry: Semester 2
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PLAY. Terms in this set
(...) Double

Replacement

Precipitation Reactions.

The positive ions in two
compounds switch

places. A precipitation
reaction will produce

an insoluble salt as one
of the reactants. Mole.

Mole: A unit of

measurement used in
chemistry to express

amounts of a chemical
substance, defined as

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the amount of any substance that contains as atoms, molecules, ions, or electrons as there are atoms in 12 grams of pure C^{12} , the isotope of ...

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(and Electrochemistry)

Practice Problems

(Skills) Search. Create.

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Guide Learning goals

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Checklist As you prepare for your final exam use the following list to monitor your progress. Only check off a learning goal when you are confident of your ability to complete the task described. Unit 0: Being a Chemist Recognize and apply safe practices when working in the lab.

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**Exam Practice Test
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Honour Chemistry:
Practice Final Exam
(Semester 2) Part A:
Multiple Choice and
Numerical Response 1.
Which of the following
equations is associated
with the largest energy
change per mole of
fluorine? A. $F_2(g) \rightarrow F_2(l)$
B. $19\ 9\ F + 4\ 2\ He \rightarrow$
 $1\ 0\ n + 22\ 11\ Na$ C. 2
 $F_2(g) + 2\ H_2O(l) \rightarrow O_2$
 $(g) + 4\ HF(aq)$

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**Final Exam Practice
Test (Semester 2) -
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Chemistry Semester 2:
Final Exam The Final
Exam in this class is
the Michigan
Secondary Credit
Examination for
Chemistry. Exam
Grades will be based
on achieving the
Department of
Education
recommended passing
rate of 70% (For
Example - correctly

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answering 70% of the questions = 100% recorded in Powerschool as your Exam score)

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includes: Stoichiometry
Gas Laws Solutions
Ions in Aqueous
Solutions Acids and
Bases Oxidation and
Reduction Reactions

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(and Electrochemistry)
Practice Problems
(Skills) ...

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including all of the
information and
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17 Formulae and
Constants L, mL =
liters, milliliters K = °C
+ 273 g = grams Gas

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constant $R = 8314$
 $\text{J/molK nm} =$
nanometers = 00821
 $\text{Latm/molK atm} =$
atmospheres J, kJ =
joules, kilojoules
Standard T and P =
273K and 10 atm V =
volts mol = moles N_A
= $602 \times 10^{23} \text{ N}$

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2 Honors chemistry
semester 2 final exam.
Which of the following

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molecular properties is
a main component of
the chemical potential
energy of matter? A.
Vibrational motion C.
Movement from place
to place. B Honors
chemistry semester 2
final exam.
Intramolecular bonding
D. Rotation about the
molecules' centre of
mass.

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