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Chemistry Notes

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Hess's Law, Entropy

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**Physical Chemistry-  
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various kinds of energy  
in physical and  
chemical systems.

Thermodynamics 1.

Predicts the feasibility  
of a physical process or  
chemical reaction  
under given condition  
of temperature and

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Chemical Thermodynamics  
pressure. 2. Predicts whether a chemical reaction would occur spontaneously or not under a given set of conditions 3.

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Chemical

Thermodynamics.

Chemical reactions, such as those that occur when you light a match, involve changes in energy as

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well as matter.

Societies at all levels of development could not function without the energy released by chemical reactions.

Almost 85% of energy consumption comes from the combustion of petroleum products, coal, wood, and garbage..We use this energy to produce electricity; to transport food, raw materials, manufactured goods; for industrial



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**Thermodynamics -**

**Read Definition,**

**Notes, Formula ...**

JEE Main Chemical

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Download. A branch of

Physics that deals with

heat, work, and

temperature, and their

relation to energy,

radiation, and physical

properties of matter is

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Thermodynamics. The four laws of Thermodynamics govern the behaviour of these quantities.

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Chemical Engineering Thermodynamics Course Notes. This note covers the following topics: Thermodynamic Property Relationships,

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Phase Equilibria,  
Principle of Phase  
Equilibrium, Applied  
Phase Equilibrium,  
Chemical Equilibrium ,  
Solving Algebraic  
Equations, Process  
Simulator. Author(s):  
T.K. Nguyen, Chemical  
and Materials  
Engineering, Cal Poly  
Pomona

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Thermodynamics is the study of the energy, principally heat energy, that accompanies chemical or physical changes. Some chemical reactions release heat energy; they are called exothermic reactions, and they have a negative enthalpy change. Others absorb heat energy and are called endothermic reactions, and they have a positive

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enthalpy change.

Chemical  
**Introduction to  
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Thermodynamic relationships determine the equilibrium composition of product mixtures resulting from chemical reaction. The nature of the product species, their proportions, their temperature, and their pressure depend upon

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**The Chemistry of  
Combustion |  
Thermodynamics**

Notes for Chemical  
Equilibrium

(Thermodynamics)

Chemical Equilibrium.

Understanding how to

work problems for

chemical equilibrium is

one of the most.

important concepts in

second-semester

chemistry. In fact, we

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will spend much of the rest of the course discussing aspects of this concept.

## **Notes for Chemical Equilibrium (Thermodynamics)**

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Spontaneous process:

The irreversible

process which

proceeds by itself is

called spontaneous

process. A spontaneous

process always tries to

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... First Law of

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internal energy (Opens  
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internal energy and  
work example (Opens a  
modal) Heat and  
temperature

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thermodynamics, we  
can predict the amount  
of energy needed to  
change a system from  
an equilibrium state to  
another. For example it  
will take about 75 kJ to  
change 1 kg of air at

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25oC and 1 atm to  
100oC and 1 atm. It  
will take much more  
energy, about 2257 kJ,  
to change

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Thermodynamics II**

This section provides  
the lecture notes for  
the course along with  
the schedule of lecture  
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Kinetics » Lecture  
Notes... chemical  
potential : 15:  
Thermodynamics  
Chemical equilibrium :  
16: Temperature,  
pressure and  $K_p$

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Thermodynamics &  
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...

A system which can exchange both energy and matter with its surroundings. A system which permits passage of energy but not

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mass, across its boundary. A system which can neither exchange energy nor matter with its surrounding. Part of the universe other than system, which can interact with it.

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Lecture notes in  
KP8108 Advanced  
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the 2nd law, modes of  
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Section 10

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Process Plant Schema

Chapter 2: Volumetric

Properties of Real

Fluids Section 1 :

General P-V-T

Behaviour of Real

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